

CHEMISTRY (CHEM)**CLASS - XI**

Full Marks 100

THEORY - 70 Marks

		Marks
Unit – I	Some Basic Concepts of chemistry	03
Unit – II	Structure of Atom	06
Unit – III	Classification of Elements and Periodicity in properties	04
Unit- IV	Chemical bonding and Molecular Structure	05
Unit – V	State of Matter; Gases and Liquids	04
Unit- VI	Thermodynamics	06
Unit- VII	Equilibrium	06
Unit- VIII	Redox Reactions	03
Unit- IX	Hydrogen	03
Unit- X	s-Block Elements	05
Unit- XI	Some p-Block Elements	07
Unit- XII	Organic Chemistry: some basic Principles and Techniques	07
Unit- XIII	Hydrocarbons	08
Unit- XIV	Environmental Chemistry	03
Total-		70

Unit – I: Some Basic Concepts of chemistry**(periods 12)**

General Introduction: Importance and scope of chemistry.

Historical approach to particulate nature of matter, laws of chemical combination. Dalton's atomic theory: concept of elements, atoms and molecules.

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Atomic and molecular masses. Mole concept and molar mass: percentage composition, empirical and molecular formula; chemical reactions, stoichiometry and calculations based on stoichiometry.

Unit – II: Structure of atoms (periods 16)

Discovery of electrons, proton and neutron; atomic number, isotopes and isobars. Rutherford's model and its limitations. Bohr's model and its limitations, concept of shell and sub shells, dual nature of matter and light, De Broglie's relationship. Heisenberg uncertainty principle, concept of orbitals, quantum numbers, shapes of s, p, and d orbitals, rules for filling electrons in orbitals - Aufbau principle, Pauli exclusion principle and Hund's rule, electronic configuration of atoms, stability of half filled, completely filled orbitals.

Unit – III: Classification of elements and Periodicity in Properties (periods 08)

Significance of classification, brief history of the development of periodic table. Modern periodic law and the present form of periodic table, periodic trends in properties of elements – atomic radii, ionic radii, ionization enthalpy, electron gain enthalpy, electronegativity, valency, nomenclature of elements with atomic number greater than 100.

Unit – IV: Chemical Bonding and Molecular Structure (periods 16)

Valence electrons, ionic bond, bond parameters, covalent bond: Born Haber Cycle. Lewis structure, polar character of covalent bond, covalent character of ionic bond, valence bond theory, resonance, geometry of covalent molecules. VSEPR theory, concept of hybridization, involving s, p and d orbitals and shapes of some simple molecules, Molecular orbital theory of homonuclear diatomic molecules and hydrogen bond.

Unit – V: States Of Matter: Gases and Liquids (periods 14)

Three states of matter. Intermolecular interactions, types of bonding, melting and boiling points. Role of gas laws in elucidating the concept of the molecule. Boyle's law, Charles' law, Gay Lussac's Law, Avogadro's Law, Ideal Behaviour, empirical derivation of gas equation. Avogadro's number, Ideal gas equation. Derivation from ideal behaviour, Liquefaction of gases, critical temperature, kinetic energy and molecular speeds (elementary idea)
Liquid state – vapour pressure, viscosity and surface tension (qualitative idea only, no mathematical derivations).

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Unit – VI: Chemical Thermodynamics (periods 16)

Concepts of system, types of systems, surroundings. Work, heat, energy, extensive and intensive properties, state functions.

First law of thermodynamics – internal energy change (ΔU) and enthalpy change (ΔH). Hess's law of constant heat summation, enthalpy of bond dissociation, combustion, formation, atomization, sublimation, Phase transformation, ionization, and solution.

Introduction of entropy as a state function, Gibbs energy change for spontaneous and non spontaneous processes, criteria for equilibrium.

Second and third laws of thermodynamics.

Unit – VII: Equilibrium (periods 20)

Equilibrium in physical and chemical processes, dynamic nature of equilibrium, law of mass action, equilibrium constant, factors affecting equilibrium – Le chatelier's principle; ionic equilibrium – ionization of acids and bases, strong and weak electrolytes, degree of ionization of polybasic acids, acid strength, concept of pH Henderson Equation. Hydrolysis of salts (elementary idea). Buffer solutions, solubility product, common ion effect (with illustrative examples).

Unit – VIII: Red ox Reactions (periods 06)

Concept of oxidation and reduction, red ox reactions, oxidation number, balancing redox reactions in terms of loss and gain of electrons and change in oxidation number.

Unit – IX: Hydrogen (periods 08)

Position of hydrogen in periodic table, occurrence, isotopes, preparation, properties and uses of hydrogen; hydrides – ionic, covalent and interstitial; physical and chemical properties of water, heavy water; hydrogen peroxide-preparation, properties, structure and use; hydrogen as a fuel.

Unit – X: s-Block Elements (alkali and Alkaline earth metals) (periods 16)

Group 1 and Group 2 elements:

General introduction, electronic configuration, occurrence, anomalous properties of the first element of each group, diagonal relationship, trends in the variation of properties (such as ionization enthalpy, atomic and ionic radii), trends in chemical reactivity with oxygen, water, hydrogen and halogens; uses.

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preparation and properties of some important compounds:

Sodium carbonate, sodium hydroxide and sodium hydrogen carbonate, biological importance of sodium and potassium.

CaO, CaCO₃ and industrial use of lime and limestone, biological importance of Mg and Ca

Unit –X I: Some p-Block Elements

(periods 14)

General Introduction to p-Block Elements

Group 13 elements: General introduction, electronic configuration, occurrence. Variation of properties, oxidation states, trends in chemical reactivity, anomalous properties of first element of the group; Boron – physical and chemical properties, some important compounds: borax, boric acid, boron hydrides, Aluminium: reactions with acids and alkalis and uses.

Group 14 elements: General introduction, electronic configuration, occurrence, variation of properties, oxidation state, trends in chemical reactivity, anomalous behaviour of first element, carbon- catenation, allotropic forms, physical and chemical properties; uses of some important compounds; oxides.

Important compounds of silicon and a few uses: silicon tetrachloride, silicones, silicates and zeolites, their uses and structure of silicates.

Unit –XII: Organic chemistry – Some Basic Principles and Techniques

(periods 16)

General introduction, methods of qualitative and quantitative analysis, classification and IUPAC nomenclature of organic compounds

Electronic displacements in a covalent bond: inductive effect, electrometric effect, resonance and hyper conjugation.

Homolytic and Heterolytic fission of a covalent bond: free radicals, carbocations, carbanions, electrophiles and nucleophiles, types of organic reactions.

Unit –XIII: Hydrocarbons

(periods 16)

Classification of hydrocarbons

Alkanes – Nomenclature, isomerism, conformations (ethane only), physical properties, chemical reactions including halogenations, free radical mechanism, combustion and pyrolysis.

Alkenes – Nomenclature, structure of double bond (ethene), geometrical isomerism, physical properties, methods of preparation; chemical reactions; addition of hydrogen, halogen, water,

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hydrogen halides (markovnikov's addition and peroxide effect), ozonolysis, oxidation, mechanism of electrophilic addition.

Alkynes – Nomenclature, structure of triple bond (ethyne), physical properties. Methods of preparation, chemical reactions; acidic character of Alkynes, addition reaction of – hydrogen, halogens, hydrogen halides and water.

Aromatic hydrocarbons; Introduction, IUPAC nomenclature; Benzene; resonance aromaticity; chemical properties; mechanism of electrophilic substitution – nitration, sulphonation, halogenation, Friedel craft's alkylation and acylation, carcinogenicity and toxicity.

Unit –XIV: Environmental chemistry

(periods 08)

Environmental pollution – air, water and soil pollution, chemical reactions in atmosphere, smog, major atmospheric pollutants; acid rain, ozone and its reactions, effects of depletion of ozone layer, greenhouse effect and global warming – pollution due to industrial wastes; green chemistry as an alternative tool for reducing pollution, strategy for control of environmental pollution.

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Practical

Marks - 30

Evaluation Scheme for Examination

	Marks
Volumetric analysis	10
Salt Analysis	08
Content Based Experiment	06
Class Record and Viva	06
Total	30

Practical Syllabus

Total Periods 60

A. Basic Laboratory Techniques

(periods 03)

- I. Cutting glass tube and glass rod
- ✓ II. Bending a glass tube
- III. Drawing out a glass jet
- IV. Boring a cork

B. Characterization and purification of chemical substances

(periods 07)

- ✓ I. Determination of melting point of an organic compound
- ✓ II. Determination of boiling point of an organic compound
- III. Crystallization of impure sample of anyone of the following: Alum, copper sulphate, Benzoic acid.

C. Experiments related to pH change

a. Anyone of the following experiments:

- ✓ Determination of pH of some solutions obtained from fruit juices, varied concentrations of acids, bases and salts using pH paper or universal indicator.

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- ✓ Comparing the pH of solutions of strong and weak acid of same concentration.
- ✓ Study the pH change in the titration of a strong base using universal indicator.

b. Study of pH change by common-ion effect in case of weak acids and weak bases.

D. Chemical equilibrium

(periods 05)

One of the following experiments:

- a) Study the shift in equilibrium between ferric ions and thiocyanate ions by increasing/decreasing the concentration of either ions.
- b) Study the shift in equilibrium between $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ and chloride ions by changing the concentration of either of the ions.

E. Quantitative estimation

(periods 18)

- Using a chemical balance.
- Preparation of standard solution of oxalic acid.
- Determination of strength of a given solution of sodium hydroxide by titrating it against standard solution of oxalic acid.
- Preparation of standard solution of sodium carbonate.
- Determination of strength of a given solution of hydrochloric acid by titrating it against standard sodium carbonate solution.

F. Qualitative analysis

Determination of one anion and one cation in a given salt

Cautions:- Pb^{2+} , Cu^{2+} , As^{3+} , Al^{3+} , Fe^{3+} , Mn^{2+} , Ni^{2+} , Zn^{2+} , Co^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , Mg^{2+} , NH_4^+

Anions:- CO_3^{2-} , S^{2-} , SO_3^{2-} , SO_4^{2-} , NO_2^- , NO_3^- , Cl^- , Br^- , I^- , PO_4^{3-} , $\text{C}_2\text{O}_4^{2-}$, CH_3COO^-

(Note: Insoluble salts excluded)

G. Detection of nitrogen, sulphur, chlorine

(periods 12)

CHEMISTRY (CHEM)**QUESTION PATTERN****Class XI (Theory)****Time: 3 Hours****Marks: 70**

Sl.no	UNIT	VSA (1Marks)	SA I (2 Marks)	SA II (3 Marks)	LA (5Marks)	Total
1	Some Basic Concepts of Chemistry	1 (1)	2 (1)	-	-	03
2	Structure of Atom	-	-	6 (2)	-	06
3	Classification of Elements and Periodicity in Properties	-	4 (2)	-	-	04
4	Chemical Bonding & molecular structure	2(2)		3(1)	-	05
5	States of Matter: Gases & Liquids	-	4(2)	-	-	04
6	Thermodynamics	-	-	6 (2)	-	06
7	Equilibrium	2(2)	4(2)	-	-	06
8	Redox Reactions	-	-	3 (1)	-	03
9	Hydrogen	1(1)	2(1)	-	-	03
10	s-Block Elements	-	-	-	5 (1)	05
11	Some p-Block Elements	-	2(1)	-	5 (1)	07
12	Organic Chemistry— Some basic principles & techniques		2(1)	-	5 (1)	07
13	Hydrocarbons			3(1)	5(1)	08
14	Environmental Chemistry			3 (1)	-	03
	Total	6 (6)	20 (10)	24 (8)	20 (4)	70

CHEMISTRY (CHEM)**CLASS - XII**

Full Marks 100

THEORY - 70 Marks

	Marks
Unit – I Solid State	04
Unit – II Solutions	05
Unit – III Electrochemistry	05
Unit- IV Chemical Kinetics	05
Unit – V Surface Chemistry	04
Unit- VI General principles and processes of Isolation of Elements	03
Unit- VII p-Block Elements	08
Unit- VIII d-and f- Block Elements	05
Unit- IX Coordination Compounds	03
Unit- X Haloalkanes and Haloarenes	04
Unit- XI Alcohols, Phenols and Ethers	04
Unit- XII Aldehydes, Ketones and Carboxylic acids	06
Unit- XIII Organic Compounds containing Nitrogen	04
Unit- XIV Bio molecules	04
Unit-XV Polymers	03
Unit-XVI Chemistry in Everyday Life	03

Total- 70

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Unit – I: Solid State

(periods 12)

Classification of solids based on different binding forces: molecular, ionic, covalent and metallic solid, amorphous and crystalline solids (elementary idea), unit cell in two dimensional and three dimensional lattices, packing efficiency, calculation of density of unit cell, packing in solids, voids, number of atoms per unit cell in a cubic unit cell, point defects, electrical and magnetic properties. Band theory of metals, conductors, semiconductors and insulators and n & p type semiconductors.

Unit – II: Solutions

(periods 12)

Types of solutions, expression of concentration of solution of solids in liquids, solubility of gases in liquids, solid solutions, colligative properties – relative lowering of vapour pressure, Raoult's law, elevation of boiling point, depression of freezing point, osmotic pressure, determination of molecular masses using colligative properties, abnormal molecular mass, van't Hoff factor and calculations involving it.

Unit – III: Electrochemistry

(periods 14)

Red ox reactions, conductance in electrolytic solutions, specific and molar conductivity, variations of conductivity with concentration, Kohlrausch's law, electrolysis and laws of electrolysis (elementary idea), dry cell – electrolytic cells and Galvanic cells; lead accumulator, EMF of a cell, standard electrode potential, Nernst equation and its application to chemical cells. Reaction between Gibbs energy change and emf of a cell, fuel cells: corrosion.

Unit – IV: Chemical Kinetics

(periods 12)

Rate of a reaction (average and instantaneous), factors affecting rates of reactions; concentration, temperature, catalyst; order and molecularity of a reaction; rate law and specific rate constant, integrated rate equations and half life (only for zero and first order reactions); concept of collision theory (elementary idea, no mathematical treatment), activation energy, Arrhenius equation.

Unit – V: Surface Chemistry

(periods 08)

Adsorption – physisorption and chemisorption; factors affecting adsorption of gases on solids; catalysis; homogenous and heterogeneous, activity and selectivity; enzyme catalysis; colloidal state, distinction between true solutions, colloids and suspensions; lyophilic, lyophobic, multimolecular colloids; properties of colloids; Tyndall effect, Brownian movement, electrophoresis, coagulation; emulsion – types of emulsions. Elementary idea of nanomaterials.

Unit – VI: General Principles and Processes of Isolation of Elements (periods 08)

Principles and methods of extraction – concentration, oxidation, reduction electrolysis methods and refining; occurrence and principles of extraction of aluminium, copper, zinc and iron.

Unit – VII: p- Block Elements (periods 14)

Group 15 elements: general introduction, electronic configuration, occurrence, oxidation states, trends in physical and chemical properties; nitrogen – preparation, properties and uses; compounds of nitrogen; preparation and properties of ammonia and nitric acid, oxides of nitrogen (structure only); Phosphorus – allotropic forms; compounds of phosphorus; preparation and properties of phosphine, halides (PCl_3 , PCl_5) and oxoacids (elementary idea only).

Group 16 elements: General introduction, electronic configuration, oxidation states, occurrence, trends in physical and chemical properties; dioxygen; preparation, properties and uses; classification of oxides; ozone, sulphur – allotropic forms;

Compounds of sulphur dioxide; sulphuric acid: industrial process of manufacture, properties and uses, other oxides and oxoacids of sulphur (structures only).

Group 17 elements: General introduction, electronic configuration, oxidation states, occurrence, trends in physical and chemical properties; compounds of halogens; preparation, properties and uses of chlorine and hydrochloric acid, interhalogen compounds, oxoacids of halogens (structure only).

Group 18 elements: General introduction, electronic configuration. Occurrence, trends in physical and chemical properties uses.

Unit – VIII: d and f Block Elements (periods 14)

General introduction electronic configuration, occurrence and characteristics of transition metals, general trends in properties of the first row transition metals – metallic character, ionization enthalpy, oxidation states, ionic radii, colour, catalytic property, magnetic properties, interstitial compounds, alloy formation. Preparation and properties of $\text{K}_2\text{Cr}_2\text{O}_7$ and KMnO_4 .

Lanthanoids – electronic configuration, oxidation states, chemical reactivity and lanthanoid contraction and its consequences.

Actinoids - Electronic configuration, oxidation states and comparison with lanthanoids.

Unit – IX: Coordination Compounds (periods 12)

Coordination compounds – Introduction, Ligands, coordination number, colour, magnetic properties and shape, IUPAC nomenclature of mononuclear coordination compounds. Bonding (Werner's

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theory, VBT and CFT); structural and stereo isomerism, importance of coordination compounds (in qualitative inclusion of analysis, extraction of metals and biological systems)

Unit – X: Haloalkanes and Haloarenes

(periods 12)

Haloalkanes:

Nomenclature, nature of C-X bond, physical and chemical properties, mechanism of substitution reactions. Stability of carbonations, R-S and d-l configurations

Haloarenes:

Nature of C-X bond, substitutions reactions (directive influence of halogen for monosubstituted compound only, stability of carbocations R-S and d-l configurations)

Uses and environmental effects of – dichloromethane, trichloromethane, tetrachloromethane, iododorm, freons, DDT.

Unit – XI: Alcohols, phenols and Ethers

(periods 12)

Alcohols: Nomenclature, methods of preparation, physical and chemical properties (of primary alcohols only); identification of primary, secondary and tertiary alcohols; mechanism of dehydration, uses of methanol and ethanol.

Phenols: Nomenclature, methods of preparation, physical and chemical properties, acidic nature of phenol, electrophilic substitution reactions, uses of phenol.

Ethers: Nomenclature, methods of preparation, physical and chemical properties, uses.

Unit – XII: Aldehydes, Ketones and Carboxylic Acids

(periods 12)

Aldehydes and Ketones: Nomenclature, nature of carbonyl group, methods of preparation, physical and chemical properties, mechanism of nucleophilic addition, reactivity of alpha hydrogen in aldehydes; uses.

Carboxylic Acids: Nomenclature, acidic nature, methods of preparation, physical and chemical properties, uses.

Unit – XIII: Organic compounds containing Nitrogen

(periods 10)

Nitro compounds: General methods of preparation and chemical reactions.

Amines: Nomenclature, classification, structure, methods of preparation, physical and chemical properties, uses, identification of primary, secondary and tertiary amines.

Cyanides and Isocyanides – will be mentioned at relevant places in context.

Diazonium salts: Preparation, chemical reactions and importance in synthetic organic chemistry.

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Unit – XIV: Bio molecules

(periods 12)

Carbohydrates - Classification (aldoses and Ketoses), monosaccharides (glucose and fructose), D-L configuration, oligosaccharides (sucrose, lactose, maltose), polysaccharides (starch, cellulose, glycogen); importance.

Proteins - Elementary idea of α - amino acids, peptide bond, polypeptides, proteins, primary structure, secondary structure, tertiary structure and quaternary structure (qualitative idea only), denaturation of proteins; enzymes.

Lipids and hormones, their classification and functions.

Vitamins: Classification and functions.

Nucleic Acids: DNA & RNA.

Unit – XV: Polymers

(periods 08)

Classification – natural and synthetic, methods of polymerization (addition and Condensation), copolymerization. Some important polymers: natural and synthetic like poly, nylon, polyesters, Bakelite, rubber, biodegradable and non-biodegradable polymers.

Unit – XVI: Chemistry in Everyday Life

(periods 08)

1. **Chemicals in medicines**- analgesics, tranquilizers, antiseptics, disinfectants, antimicrobials, anti fertility drugs, antibiotics, antacids, antihistamines, antioxidants.
2. **Chemicals in food** – preservatives, artificial, sweetening agents.
3. **Cleansing agents** – soaps and detergents, cleansing action.

SYLLABUS

Practical

Marks 30

Evaluation Scheme for Examination

	Marks
Volumetric analysis	10
Salt Analysis	08
Content Based Experiment	06
Class Record ,Viva and Project work	06
Total	30

Practical Syllabus

A. Surface Chemistry.

- Preparation one lyophilic and one lyophobic sol. Lyophilic sol – starch albumin and gum, Lyophobic sol – aluminium hydroxide, ferric hydroxide, arsenious sulphide.
- Study of the role of emulsifying agents in stabilizing the emulsions of different oils.

B. Chemical kinetics

- Effect of concentration and temperature on the rate of reaction between sodium thiosulphate and hydrochloric acid.
- Study of reaction rates of any one of the following:
 - Reaction of iodide ion with hydrogen peroxide at room temperature using different concentrations of iodide ions.
 - Reaction between potassium iodide, KIO_3 and sodium sulphate: ($\text{Na}_2\text{S}_2\text{O}_3$) using starch solution as indicator (clock reaction).

C. Thermo chemistry

Any of the following experiments

- Enthalpy of dissolution copper sulphate or potassium nitrate.
- Enthalpy of neutralization of strong acid (HCl) and strong base (NaOH)
- Determination of enthalpy change during interaction (hydrogen bond formation) between acetone and chloroform

D. Electro chemistry

Variation of cell potential in $\text{Zn}/\text{Zn}^{2+}/\text{Cu}^{2+}/\text{Cu}$ with change in concentration of electrolytes (CuSO_4 or ZnSO_4) at room temperature.

E. Chromatography

- Separation of pigments from extracts of leaves and flowers by paper chromatography and determination of Rf Values.
- Separation of constituents present in an inorganic mixture containing two cations only (constituents having large difference in Rf values to be provided).

F. Preparation of Inorganic Components

- Preparation of double salt of ferrous ammonium sulphate or potash alum
- Preparation of potassium ferric oxalate

G. Preparation of Organic compounds

Preparation of any two of the following compounds

- Acetanilide
- Di benzal acetone
- p-Nitroacetanilide
- Aniline yellow or 2- Napthal aniline dye.
- Iodoform

H. Tests for the functional groups present in organic compounds:

Unsaturation, alcoholic, phenolic, aldehydic, ketonic, carboxylic and amino (primary) groups.

I. Characteristic test of carbohydrates, fats and proteins in pure samples and their detection in given food stuffs.

J. Determination of concentration/ minority of KMnO_4 solution by titrating it against a standard solution of:

- Oxalic acid
- Ferrous ammonium sulphate

(students will be required to prepare standard solution by weighing themselves)

K. Qualitative analysis

Determination of one cation and one anion in a given salt.

Cations – Pb^{2+} , Cu^{2+} , As^{3+} , Al^{3+} , Fe^{3+} , Mn^{2+} , Ni^{2+} , Zn^{2+} , Co^{2+} , Ca^{2+} , Sr^{2+} , Ba^{2+} , Mg^{2+} , NH_4^+

Anions – CO_3^{2-} , S^{2-} , SO_3^{2-} , NO_2^- , NO_3^- , Cl^- , Br^- , I^- , PO_4^{3-} , $\text{C}_2\text{O}_4^{2-}$, CH_3COO^-

(Note; Insoluble salts excluded)

Project work- where feasible may include

- Model preparation
- Investigatory project
- Science exhibits
- Participation in science fairs
- Testing of purity of food articles like butter, pulse and milk etc.

CHEMISTRY (CHEM)**QUESTION PATTERN****Class XII (Theory)****Time Allowed: 3Hrs****Maximum Marks: 70**

Sl.no	UNIT	VSA (1)	SAI(2)	SA II(3)	LA (5)	TOTAL
1	Solid State		4 (2)			4 (2)
2	Solutions				5 (1)	5 (1)
3	Electro Chemistry		2 (1)	3 (1)		5 (2)
4	Chemical Kinetics	1 (1)	4 (2)			5 (3)
5	Surface Chemistry	1 (1)		3 (1)		4(2)
6	General Principles and Processes			3 (1)		3 (1)
7	p-Block Elements	1 (1)	4 (2)	3 (1)		8 (4)
8	d & f- Block Elements				5 (1)	5 (1)
9	Co-ordination Compounds	1 (1)	2 (1)			3 (2)
10	Haloalkanes and Haloarenes	1 (1)		3 (1)		4 (2)
11	Alcohols, Phenols & Ethers	1 (1)		3 (1)		4 (2)
12	Aldehydes, Ketones & Carboxylic Acids	1 (1)			5 (1)	6 (2)
13	Organic Compounds containing Nitrogen		4 (2)			4 (2)
14	Bio molecules	1 (1)		3 (1)		4 (2)
15	Polymers			3 (1)		3 (1)
16	Chemistry in Everyday Life			3 (1)		3 (1)
	TOTAL:	8 (8)	20 (10)	27 (9)	15 (3)	70 (30)

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DESIGN

Sl.no	Type of Question	Marks for each question	No. of question	Total Marks
1	Long Answer (LA)	5	3	15
2	Short Answer- II (SA II)	3	9	27
3	Short Answer- I (SA-I)	2	10	20
4	Very Short Answer (VSA)	1	8	8
	Total		30	70